Lesson 4 - Classifications

Learning Intentions

By the end of this lesson I will be able to:

Explain how the elements are organised in the periodic table.

Name three groups of the periodic table and describe their reactivity.

Classification of Elements

Although the periodic table may seem confusing and organised at random when we first look at it, The Periodic Table of Elements is not just a random jumble of letters and numbers at all. It has a very definite order and structure that is related to the properties of the elements within it.

The **physical properties** of an element or substance are related to what it looks like. For example we might say 2Bromine is a brown liquid." Or "Aluminium is a shiny grey solid.".

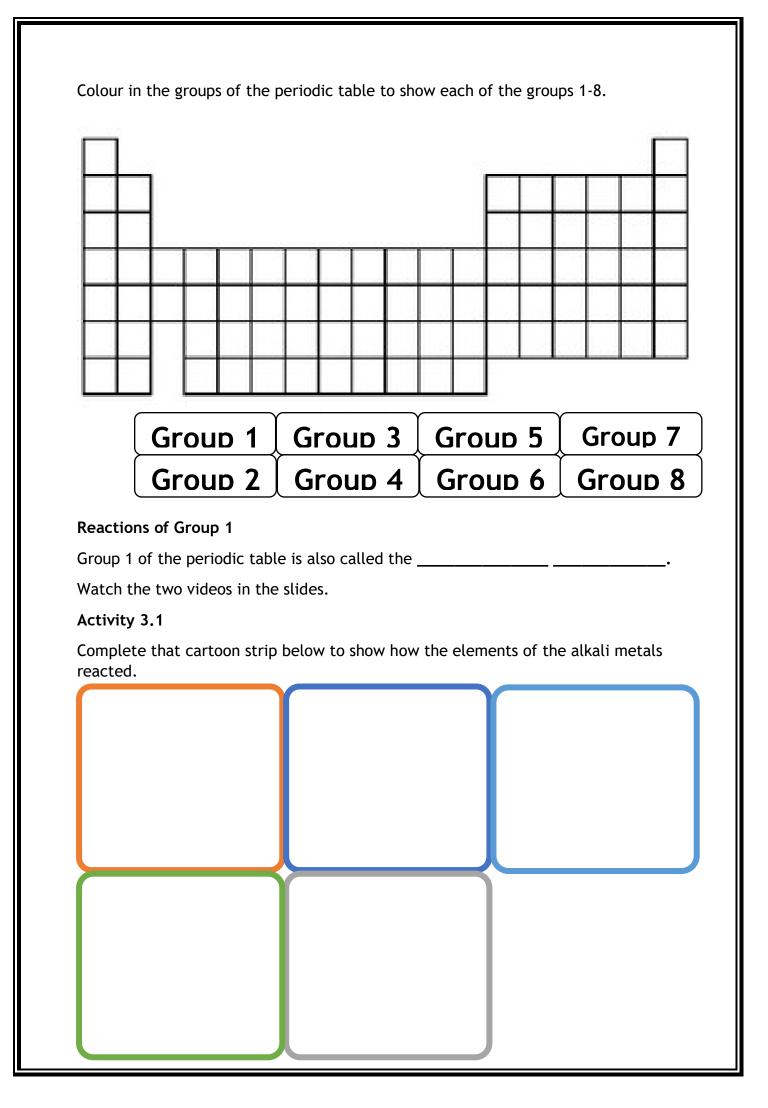
The **chemical properties** of a substance are related to how the substance behaves during reactions. For example, we could say "it burns with a squeaky pop" or "it burns bright white".

The periodic table is arranged in two basic ways, **groups** which run vertically (up and down the way) and periods which run horizontally (across the way).

Groups of the Periodic Table

The Periodic Table was first put together by Dimitri Mendeleev. He noticed that some elements looked alike and behaved in the same way, much like we may do with our family. For this reason he thought they may belong to the same' family' - or group - of elements.

The groups of the periodic table run vertically.



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3	(iive a	use for	Neon

4. Comment of the flammability of the noble gases.

5. What does krypton mean?

6. State a fact about xenon and radon.

Was I successful?	Red	Yellow	Green
I can explain how the elements are organised in the Periodic Table.			
I can name 3 groups of the periodic table and describe their reactivity.			